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CONSTANTS . Throughout the
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have the definitions

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2COOH(aq) + H₂O(l) CH₃CH

2COO⁻(aq) + H₃O⁺(aq) 2.

Propanoic acid, CH₃CH

2COOH, is a carboxylic acid

that reacts with water

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(d) Calculate the number of moles of precipitate that is produced in the experiment.

$$1.698 \text{ g} - 1.462 \text{ g} = 0.236 \text{ g}$$

PbI₂ (s) 2 4 22 2 1 mol PbI₂

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0.236 g PbI = 5.12 10 mol
PbI

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represent guesstimate
points) Question 1 (a) (i)
 $\text{Pb}^{2+}(\text{aq}) + 2\text{I}^{-}(\text{aq}) \rightarrow \text{PbI}_2(\text{s})$
(1) (ii) Potassium and
nitrate ions are in solution

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at the beginning of the reaction and at the end, do not change in any way (i.e., they are spectators), and as such can be omitted from the equation.

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